centimeter (S.T.P.) quantities, and were measured with standard thermal conductivity equipment.

The most important general limitation on recycling indefinitely is reached when the material has spread and fills one entire section of the column. Thus, the separation of Fig. 3 could be continued until this point is reached; alternatively, the front of the peak to point A can be isolated, removed, and reinjected for a second passage. The separations reported here are typical, but optimum operating conditions for separation have not yet been systematically determined in any system. The technique has been used for 10- to 20-ml gas samples with the present equipment. JOHN W. ROOT

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## **References** and Notes

- See, for example, H. Purnell, Gas Chromatog-raphy (Wiley, New York, 1962).
   A. J. P. Martin in 1957 International Gas Chromatography Symposium (Instrument Soc. of America, Pittsburgh, 1957); in Vapour Phase Chromatography, D. H. Desty, Ed. (Putterworth London 1957) 2. A.
- *Phase Chromatography*, D. H. Desty, Ed. (Butterworth, London, 1957). K. Wilzbach and P. Riesz, *Science* **126**, 748 (1957); W. E. Falconer and R. J. Cvetanovic, 3. K

Anal. Chem. 34. 1064 (1962); R. J. Cvetanovic, F. J. Duncan, W. E. Falconer, *Canad. J. Chem.* 41, 2095 (1963); E. K. C. Lee, thesis, University of Kansas, 1963. Marked isotopic effects have also been observed in partition chroma-tography and ion exchange, but have not yet been exploited for separations of isotopically labeled molecules. See, for example, W. G. Brown, L. Kaplan, A. R. Van Dyken, K. E. Wilzbach, Proceedings of the International Conference on the Peaceful Uses of Atomic Energy, 15, 16 (United Nations, New York, 1956), and C. K. Mann and R. K. Sheline, J. Phys. Chem. 67, 1519 (1963).

- E. Gluckauf and G. P. Kitt, in *Vapour Phase* Chromatography, D. H. Desty, Ed. (Butter-worth, London, 1957); W. A. Van Hook and P. H. Emmett, J. Phys. Chem. 64, 673 (1960); F. H. Ennnett, J. Phys. Chem. 64, 675 (1960);
   W. R. Moore and H. R. Ward, J. Am. Chem. Soc. 80, 2909 (1958);
   P. Gant and K. Yang, Science 129, 1548 (1959);
   J. K. Lee, B. Mus-grave, F. S. Rowland, J. Chem. Phys. 32, 1266 (1960); (1960).
- A recycle system using 2-m packed columns has been described by M. J. E. Golay, H. I. Hill, S. D. Noren, Anal. Chem. 35, 488 (1963).
   K. Yang and P. Gant have recently separated
- all of the isotopic molecules  $CH_{4-x}D$ each other, and the molecules  $CH_{4-x}T$ one another (private communication).
- 7. This separation presumably depends upon a This separation presumably depends upon a slightly higher vapor pressure for RT than that for RH. The  $C_4D_{10}/C_4H_{10}$  separation is similarly dependent, but the  $CD_4/CH_4$  separation may involve a molecular size effect since it is affected by the hole size of the molecular sieve.
- R. Wolfgang and F. S. Rowland, Anal. Chem.
  30, 903 (1958); J. K. Lee, E. K. C. Lee,
  B. Musgrave, Y. N. Tang, J. W. Root, F. S. 8. Rowland, *ibid.* 34, 741 (1962).
- Partially supported by Air Force grant AFOSR-9. 62-15, by A.E.C. contract AT-(11-1)-407 and by predoctoral fellowship support from the Na-Foundation (J.W.R.) tional Science Foundation (J.W.R.) and the Pan American Petroleum Foundation (E.K.C.L.).

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## Dolomitization: Observations on the Island of Bonaire,

## **Netherlands Antilles**

Abstract. Evaporation of sea water and gypsum precipitation on south Bonaire produce dense brines having a high ratio of magnesium to calcium. Recent dolomite is a common replacement product in the associated carbonate sediments. A mass balance indicates that brine is being refluxed downward today, and the Pleistocene dolomite on north Bonaire could have been produced by similar reflux of brine.

The problem of dolomitization of limestones [replacement of CaCO<sub>3</sub> by  $CaMg(CO_3)_2$  has intrigued geologists for the last 100 years. Although many mechanisms have been proposed to explain dolomitization, no mechanism has yet been generally agreed upon. Observations which we recently made on the island of Bonaire in the Netherlands Antilles are consistent with the mechanism proposed to explain the dolomite found in the Permian Reef Complex (1).

In this proposed mechanism, the evaporation of sea water results in the precipitation of gypsum, which raises the ratio of magnesium to calcium of the water. The magnesium-rich water is capable of dolomitizing limestone and flows downward through the underlying sediments because it is more dense than fresh water or normal sea water. A similar mechanism has been proposed by Adams and Rhodes (2).

The south end of the island of Bon-

aire is a flat terrain close to sea level but isolated from the sea by a low coral rubble ridge. Exposed flats of soft, wet, Recent carbonate sediments cover about 34 m<sup>2</sup>, and a number of hypersaline lakes are enclosed in the area by the coral rubble ridge. The sediments are composed of gypsum, aragonite, calcite, and dolomite. Most sediments from south Bonaire contain some dolomite, and some crusts are composed of as much as 95 percent of this mineral. Radiocarbon dating of two dolomite samples from these crusts gave ages of  $1480 \pm 140$  and  $2190 \pm 150$  years. The composition of the dolomite, determined by x-ray diffraction, ranges from Mg46Ca54 to Mg44Ca56. The ordering peaks 10.1, 10.5, and 20.1 of dolomite are visible on the x-ray powder pattern. The dolomite crystal size is about 2 microns, and replacement of shells demonstrates that at least some if not all of the dolomite was formed by replacing calcium carbonate.

Chemical analyses of the water in the hypersaline lakes and from pits in the Recent sediment indicate that the water has molar magnesium/calcium ratios which average about 30/1 compared with 5/1 in sea water. The results of a number of water analyses are plotted in Fig. 1. The logarithm of the chloride concentration is plotted against the logarithm of the concentration of the other ions. The data indicate that during evaporation of the original sea water, these waters have lost calcium carbonate and calcium sulfate, but there has been no detectable loss of magnesium. The water analyses are in general agreement with data of Usiglio (in Clarke, 3) and confirm the gypsum solubility data of Posnjak (4) on gypsum.

Gypsum is a common mineral on the extensive flats of south Bonaire and is the predominant mineral in the sediments under the bottoms of the hypersaline lakes. On the other hand, the absence of halite and other late evaporite minerals in the sediments indicates

Table 1. Mass balance of Pekelmeer per unit area per year.

Source	g cm <sup>-2</sup> year <sup>-1</sup>					
	H <sub>2</sub> O	Cl-	SO <sub>4</sub> =	HCO <sub>3</sub> ~	Mg++	Ca++
Sea water influx	85	1.6	0.23	0.012	.11	0.035
Rainfall Chemical precipitates	77 —0.015		07	005*		031
Evaporation Reflux	-150 -12	-1.6	16	005* 002	11	004

\*  $2HCO_{3^-} + Ca^{++} \rightarrow CaCO_3 \downarrow + CO_2 \uparrow + H_2O \uparrow$ .

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Fig. 1. Analysis of evaporite waters. Solid lines represent simple concentration of ions with evaporation. Dashed lines show the deviation of the actual concentrations from the solid lines.

that south Bonaire is not a closed evaporating pan.

We made detailed observations on the largest of the hypersaline lakes, the Pekelmeer. We located the major zone where sea water was seeping into the lake through the coral rubble ridge and measured the rate of influx. It was sufficient to account for the difference between the rate of evaporation from the lake estimated from the evaporative pan experiments on Curaçao and the rate of rainfall given by Westermann and Zonneveld (5). The absence of halite in the lake sediments and the fact that the lake water was only about seven times the concentration of sea water indicate that there must be a continuous loss of hypersaline water from the lake. The coral rubble ridge forms a natural dike around south Bonaire which does not allow surface loss of hypersaline water. We therefore must conclude the subsurface flow (reflux) takes place through the bottom of the

lake. Reflux occurs because the density of the hypersaline water is higher than the density of the other waters present, and thus is related to evaporation. Because the interstitial water as well as the lake water is evaporating, reflux should be occurring from the entire area of south Bonaire. A mass balance of the water and the salts in the Pekelmeer allows us to estimate the rate of this reflux. The specific rates are given in Table 1.

We do not know the geometry of the refluxing waters beneath south Bonaire, nor were we able to examine the diagenesis produced by these magnesiumrich waters in traversing the older sediments. However, observations on north Bonaire suggest that Pleistocene limestones have been dolomitized by refluxing evaporite waters.

On the leeward side of north Bonaire are seaward-dipping Pleistocene or Pliocene carbonate rocks. The dip of the Pliocene-Pleistocene carbonate rocks is similar to the depositional dip of the lime sands being deposited today off the leeward coast of Bonaire, and petrographic comparison shows that the Recent sands and the Pliocene-Pleistocene rocks are very similar in sedimentary texture and fauna. We agree with Martin (6) and Schaub (7) that the dip is primary. The reconstruction of the Pliocene-Pleistocene topography of north Bonaire shown in Fig. 2 is based on the relationship between the Recent offshore sands and the Recent exposed carbonate flats. Thus, it is inferred that an area similar to south Bonaire existed above the Pliocene-Pleistocene rocks. Large dolomitized areas in the Pliocene-Pleistocene limestones cut across the bedding. The dolomite rhombs average 75 microns, and the composition determined by x-ray diffraction ranges from Mg46Ca54 to Mg44Ca56. The fact that dolomite is most abundant at the upward edge of the outcrop suggests downward movement of the dolomitizing water. The reflux of dolomitizing water from an area similar to south Bonaire would produce such a geometry in the underlying marine deposits.

The adequacy of the reflux mechanism observed on south Bonaire to explain the Pliocene-Pleistocene dolomites on north Bonaire is shown by the following calculation. The Pleistocene topography of north Bonaire could have had an evaporating surface of 10 km<sup>2</sup>. A yearly net evaporation rate of 100 cm would evaporate a total of 1010 kg

SW			NE
Reconstructed Sea Levi	el E:	xposed Flat	
Sea Level		TITHT	te the
Plio-Pleistocene Foreslope Deposits	aceous blcanic Rocks	Refluxin Dolomitizing O Me	water 300 aters

Fig. 2. Reconstruction of Pliocene-Pleistocene dolomitization.

of sea water per year. This would supply 5  $\times$  10<sup>8</sup> moles of magnesium ion per year. A rough estimate indicates that the inclined beds contain about 0.5 cubic kilometer (5  $\times$  10<sup>14</sup> cm<sup>3</sup>) of dolomite. Therefore, the time required to form the estimated volume of dolomite is of the order of 10<sup>5</sup> years.

Although these estimates do not prove that the inclined beds were dolomitized by the proposed mechanism, they indicate that the mechanism is adequate for the time available. We are not aware of any other mechanism of dolomitization that can explain qualitatively this amount and distribution of dolomite in about a million years. We believe that the conditions that occur on Bonaire may have existed many times in the past. The fine-grained Recent dolomites found on south Bonaire are similar to many ancient dolomites commonly interpreted as "primary" or "penecontemperaneous," and the Pliocene-Pleistocene dolomites from north Bonaire are very similar to the "secondary" or "replacement" dolomites found in older rocks.

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## **References and Notes**

- 1. N. D. Newel, J. K. Rigby, A. G. Fisher, A. J. Whiteman, J. E. Hickox, J. S. Bradley, The Permian Reef Complex of the Guadalupe Mountains Region, Texas, and New Mexico-A Study in Paleoecology (Freeman, San Fran-
- A Study in Functions, (1.10, 1953).
  2. J. E. Adams and M. L. Rhodes, Bull. Am. Assoc. Petrol. Geol. 44, 1912 (1960).
  3. F. W. Clarke, U.S. Geol. Surv. Bull. 770, (1997).
- (1924)
- E. Posnjak, Am. J. Sci. 238, 559 (1940).
   J. H. Westermann and J. I. S. Zonneveld, Photo-Geological Observations and Land Capabilities and Land Use Survey on the Island of Bonaire (Netherlands Antilles) (Royal Tropical
- Institute, Amsterdam, 1956). K. Martin, Bericht über eine Reise nach Niederländisch Westindien und daranf gegründate Studien (Geologisher Teil, Leiden, 1888). H. P. Schaub, Bull. Am. Assoc. Petrol. Geol. 32, 1275 (1948). 7.
- 8.
- We thank P. H. de Buisonje and the staff of Shell Curaçao for their cooperation. Present addresses: Department of Geology, University of Minnesota, Minneapolis. † Department of Oceanography, Oregon State University, Corvallis.

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