tasks (*3*), regardless of whether the injury was in the right or left hemisphere, or in the frontal, parietal, temporal, or occipital lobes. The deficit on the Army General Classification Test turned out to be comparatively focal (*1*), and so were a number of other changes in performance, such as difficulty with route finding (*4*), which was limited to the group with parietal penetration.

The work of Lashley (*5*) and others on subhuman mammals can be similarly interpreted. For certain complex tasks, such as the maze, Lashley found general (nonlocalized) effects of cortical removals in rats. With other tests, in the same animals, he found focal changes such as alterations in brightness habit after occipital removals, and difficulties on a "double platform box" after anterior removals. Thus, specific and general effects coexist after cerebral lesions in man, as well as in subhuman forms; which of these effects appears to predominate depends on the nature of the tests employed. If the range of the tasks is sufficiently extended, one finds specific and general effects in obligatory association.

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Thermal Decomposition of 8-Quinolinol Chelates of Silver (I)

The reaction of 8-quinolinol (oxine) with silver (I) has been the subject of a number of investigations. Vis (*1*) found that a green crystalline mass corresponding to the formula, $AgC_9H_6NO \cdot C_9H_6NOH$, was obtained from concentrated solutions of the reactants. Fox (*2*), however, obtained a green precipitate which corresponded to the 1:1 chelate, AgC_9H_6NO. From ionic charge considerations, one would expect the 1:1 chelate to be formed since the 8-quinolinol ion forms a five-membered ring containing one ionic and one covalent point of attachment to the metal ion.

The work of Tzinberg (*3*), which was confirmed by Hein and Regler (*4*), revealed that the formula advocated by Vis was correct. Further confirmation came from Block, Bailar, and Pearce (*5*), who were able to prepare two modifications of the chelate, a yellow form and a green one.

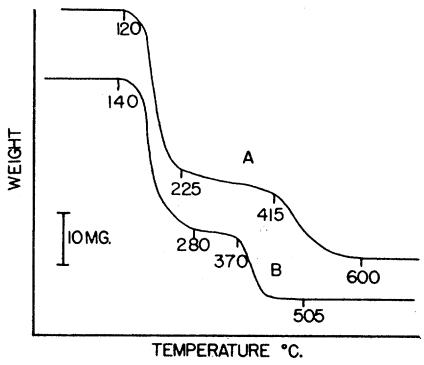

Fig. 1. Thermal decomposition curves of the 8-quinolinol chelates of silver (I). *A*, Green form; *B*, yellow form. Heating rate, 5.4°C per minute.

The nature of the bonding in the metal chelates is still open to question. It was first thought that silver (I) was oxidized to silver (II) and thus would have the normal bonding found in most of the metal 8-quinolinol chelates. If this were the case, the metal chelates would be paramagnetic; however, it was found that they were diamagnetic and hence, that they were silver (I) chelates (*5*). In view of these findings, it was suggested that the silver chelates were addition compounds containing an extra molecule of 8-quinolinol per molecule of silver chelate. This type of chelate, called a lattice chelate because the extra molecule of 8-quinolinol is thought to be held by weak lattice forces, is known for the 8-quinolinol chelates of Sc, Th, U (VI), and Pu (VI).

Now if the silver 8-quinolinol chelate contains solvated 8-quinolinol, it may be possible to prepare the normal 1:1 chelate by thermal decomposition. To find out whether this is possible, the two modifications of the silver 8-quinolinol chelate were prepared as previously described (*5*) and subjected to thermal decomposition on a thermobalance (*6*).

The thermal decomposition curves for the yellow and green forms of the silver 8-quinolinol chelate are given in Fig. 1. The yellow form was the more stable of the two modifications. It was stable up to 140°C when it began slowly to lose weight. The weight loss then became quite rapid, giving a break in the curve at 280°C; however, a constant weight level having the stoichiometry of the 1:1 chelate was not obtained. Beyond 370°C, further rapid weight loss took place to give the metallic silver level beginning at 505°C.

The green form began to lose weight at 120°C, giving a break in the curve at 225°C. Again, a horizontal weight level was not found for the 1:1 chelate. Beyond 415°C, further rapid weight loss took place to give the metallic silver level beginning at 600°C.

The results of these curves reveal that it is not possible to remove thermally the extra solvate molecule of 8-quinolinol without total disruption of the silver chelate. Such behavior is contrary to the thermal decomposition of the 8-quinolinol chelates of thorium and uranium (VI) but similar to that of scandium (*7*).

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Suggested Revision of Nomenclature—Angiotensin

Concurrent discovery has become commonplace, almost as though a mental sputnik regularly circled the earth, distributing with abandon our most exciting thoughts. The vasoactive peptide resulting from the action of renin on an alpha-globulin was thus discovered by two groups of investigators with the result that the peptide received two trivial names, angiotonin and hypertensin. Synthesis of the octapeptide has now brought a degree of certainty about the identity of this peptide and justifies dropping the double nomenclature. We propose the simplified name, *angiotensin*, and its derivatives *angiotensinase* and *angiotensinogen*. *Angiotensin* is a hybrid word but does, we think, have the advantage of being easy to pronounce even with a variety of accents, and it is euphonious and is understandable despite the most recalcitrant microphone.

There will be many who from habit will want no change, but we hope usage will make the heart grow fonder.

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