

eration of a semilogarithmic pH calculator. This yields the term $(pA - pB)$ in place of the usual $\log B/A$; in all subsequent equations or graphs the terms may be used interchangeably. The usage was retained in the second report in formulating two nomograms for glycyl aminotricarballylic acid. If the pH range of the straight lines of Fig. 1 be extended beyond four pH units, the apparent discontinuities noted by Levy do not appear. When, as in the d'Ocagne nomogram (Fig. 2), the pH scale runs from 2 to 10, no breaks occur.

The physicochemical questions concern the number of ionic species required for a given polyelectrolyte. Since each titratable group involves two forms, a proton-donor A and an acceptor B , the number of mathematically possible species s is 2^n , where n is the number of titratable groups and pK 's. For glycyl aminotricarballylic acid, n is 4 and s is 16. When n is 10, s is 1024; when n is 20, s exceeds 10^6 . Proteins may contain 100 or more titratable groups; when n is 100, s is about 10^{30} . This is more than 10^6 moles; if a molecular weight of 10^5 is assumed, the weight is about 10^8 kilograms. Obviously, only an infinitesimal fraction of the mathematically possible species can or should be considered.

The number of equations necessary to

represent a complex polyelectrolyte is much nearer to n than to s . For glycine, n is 2 and s is 4. One of these is the uncharged neutral molecule present to the extent of about one molecule in 10^8 . The curve requires two equations, relating three ionic species. When n is 4, s is 16. If these are tabulated for glycyl aminotricarballylic acid, 8 or 10 of the 16 forms are found to be of very low probable occurrence, as for example the uncharged neutral molecule. The curve may be described by four pK values relating five ionic species. For higher values of n , s becomes successively 32, 64, 128, and so on. For most purposes the distribution of charge is given by n equations and pK values, relating $(n+1)$ ionic species. A generalized nomogram is derived on this basis.

Levy, apparently well content with algebraic formulations, considers nomograms superfluous. Others, seeking elegance, find nomograms useful and rewarding. In a system containing several polyelectrolytes, algebraic formulations and curvilinear diagrams become inadequate. In biological systems there are large numbers of simultaneous reactions involving not only hydrogen ions but also other cations and anions. There are also numerous oxidation-reduction reactions which depend on pH . Algebraic

formulations consist of numerous simultaneous equations. When the number exceeds five or ten, it is difficult for one not using visual aids to coordinate all the simultaneous processes. Geometrical transformation to curved polydimensional surfaces is difficult and does little to clarify the relations. Formulation of the equations as straight lines and construction of nomograms go far toward simplifying these problems. This is a well-established procedure in many branches of science. At certain levels of complexity two-dimensional linear nomograms become preferable not only to algebraic formulations but also to curved polydimensional surfaces or their projections.

I find no statement in either of my reports asserting nonexistent advantages over standard methods. In the second, the entire emphasis was placed on the construction of a simple linear d'Ocagne nomogram illustrative of general methods for complex problems. By these methods diagrams based on three rectangular coordinates are easily transformed to nomograms with three or more parallel coordinates.

NORMAN R. JOSEPH

Department of Chemistry,
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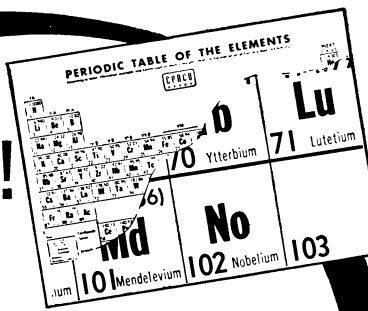
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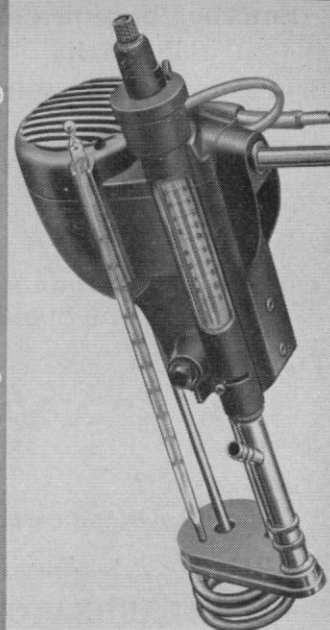
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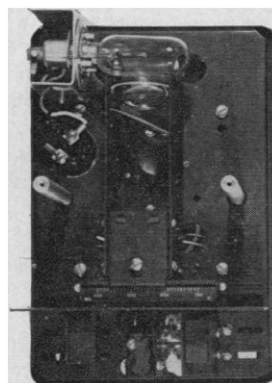
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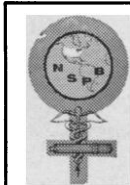
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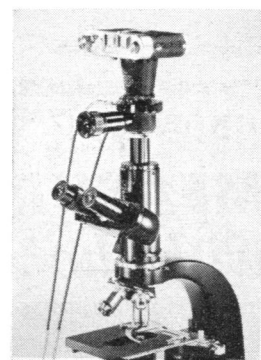
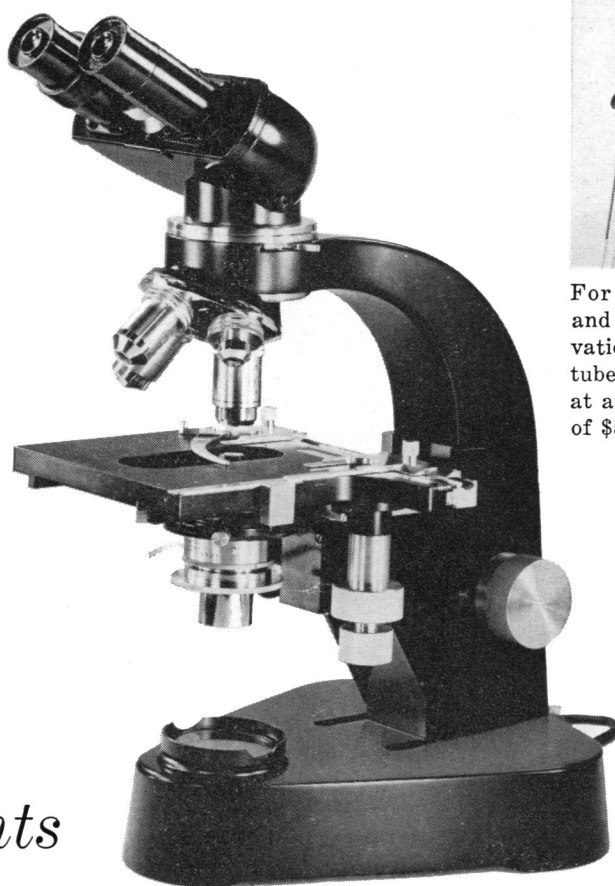
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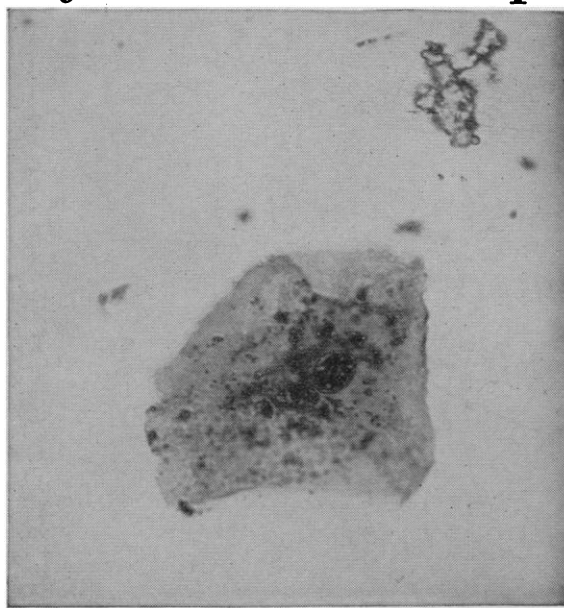


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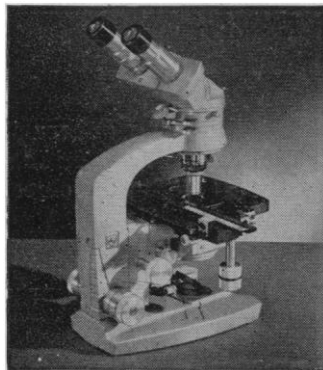
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