In my paper on "Human ecology" [*Science* **120**, 962 (1954)] the reference to "the late Otto Glaser" should read "Otto Glasser." Otto Glaser, professor emeritus of biology at Amherst, died in 1951. I am happy to report that Otto Glasser of the Department of Biophysics, Western Reserve Medical School, and editor of the excellent handbook of *Medical Physics* is still very much alive.

To O. S. Gibbs of Jefferson Medical College I am indebted for a vigorous objection to my somewhat unguarded statement that "Nature long ago discarded the nonsense of carrying poisonous wastes and nutrients in the same vessels." This is of course not literally true, except as mass effects are concerned, and they are what I had in mind. In both "pure" rivers and "pure" arterial blood there are materials with a considerable range of physiological properties, their effects being regulated by what amount to homeostatic processes. These processes break down in streams overcharged with human waste, industrial and domestic, and since we often depend on such streams for public water supply, a situation amounting to biological nonsense does exist.

Perhaps I should have used the analogy of upper and lower alimentary canal rather than artery and vein. But since the bulk of correspondence concerning the article has come from medical men, none of whom have raised the issue, I assume that my intent was clear and the analogy was not wholly inappropriate.

PAUL B. SEARS

Conservation Program,
Yale University, New Haven, Connecticut

6 January 1955.

C^{14}-Labeled Ergot Alkaloids

To provide detectable material for *in vivo* studies with the alkaloids and their derivatives and to enable an extension of our investigations on their biosynthesis, we have produced C^{14}-labeled ergot alkaloids (*1*).

Rye was grown in crocks and, prior to flowering, sealed in bell jars of 40-lit capacity. Approximately 160 ml of CO_2, containing 0.1 mc C^{14}, was introduced each day for 9 days. Illumination at an intensity of 300 ft-ca was continuous for 12 days. The plants, after removal, were infected with a culture of *Claviceps purpurea* and maintained in a normal environment until the sclerotia were mature. Fourteen sclerotia, of total weight 192 mg, were obtained.

The defatted sclerotia were extracted and the alkaloids isolated by a column partition chromatographic procedure (*2*). Further purification was effected by converting the alkaloid to the lactate and shaking with ether to remove nonalkaloid contaminants. The base, generated in aqueous solution, was recovered by shaking with ether-chloroform (2 : 1) mixture.

The ether-chloroform solution was evaporated on a planchet, and 20 1-min counts were made with a thin mica end-window G-M tube. The residue was redissolved in ether and again counted. The alkaloid was quantitatively removed, converted to the lactate, and assayed spectrophotometrically using *p*-dimethylaminobenzaldehyde test solution (U.S.P.). The results of analysis are reported in Table 1. Production of a larger quantity of labeled alkaloids is planned.

Table 1. Results of analysis.

Alkaloid	counts/min		Amt. (μg)	count/min mg	
Ergotoxine	144.8 ± 1.9	151.8 ± 3.3	68.5	2115	2216
Ergotamine	334.3 ± 4.5	358.1 ± 3.7	240.4	1390	1489
Ergonovine	42.1 ± 2.0	50.0 ± 1.6	6.0	7016	8333

ARA G. PAUL
ARTHUR E. SCHWARTING

School of Pharmacy,
University of Connecticut, Storrs

References and Notes

1. Aided by grants from the U.S. Atomic Energy Commission, contract No. AT (30-1) 1666.
2. J. E. Carless, *J. Pharm. Pharmacol.* **5**, 883 (1953).

8 December 1954.

Diffusion Constant and Diffusion Coefficient

In his treatise, *The Anatomy and Physiology of Capillaries,* Krogh (*1*, p. 268) defines the diffusion constant as

> . . . the number of cc of gas which will in one minute diffuse through an area of 1 cm^2, when the pressure gradient is one atmosphere per μ (0.001 mm).

Krogh uses this definition in his discussion of diffusion within tissues. The term *diffusion constant* has usually been regarded as analogous to the diffusivity (*2*), but Krogh's diffusion constant differs from the diffusivity, or diffusion coefficient, in two ways: in defining the gradient it employs the micron instead of the centimeter, and it employs "tension" units instead of concentration units. ("Tension" is partial pressure of diffusing substance in the gas phase at equilibrium between gas and liquid phases.) The first difference is unimportant because it simply introduces a constant factor of 10^4; but the second is important because it makes the diffusion constant a composite of two variables, the diffusivity and the solubility of diffusing substance in the liquid medium. For example, Krogh lists a value of 0.34 for the diffusion constant of O_2 in water at 20° C. The diffusivity of O_2 in water is 1.607 cm^2/day (Spoehr, *3*), or 1.12×10^{-3} cm^2/min. The solubility of O_2 in water is 0.031 at 20°C. Thus the magnitude of the diffusion constant of Krogh can be computed as $1.12 \times 10^{-3} \times 0.031 \times 10^4 = 0.346$.

Obviously this diffusion constant is not an index of diffusivity because it is so importantly influenced by solubility; yet it appears that physiologists have made the error of assuming that it is such an index.

Prosser *et al.* (*4*) reproduce Krogh's definition and