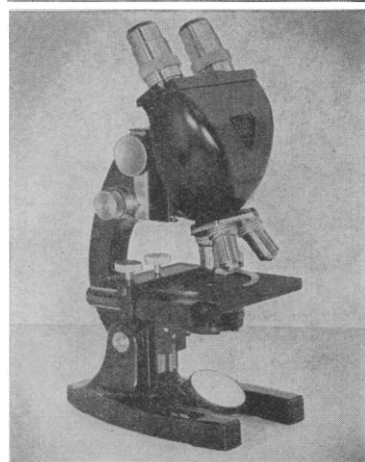
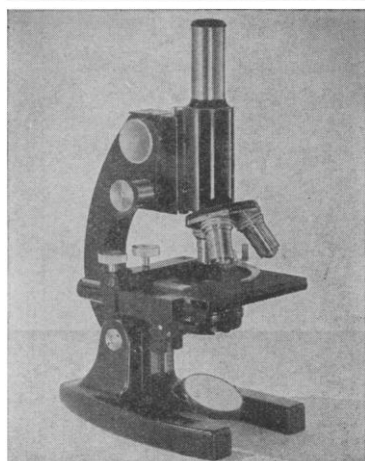


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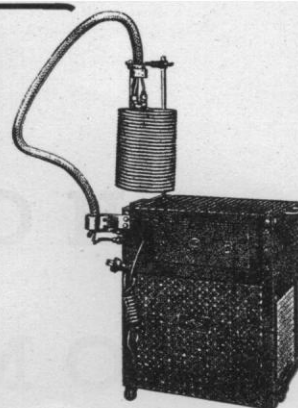
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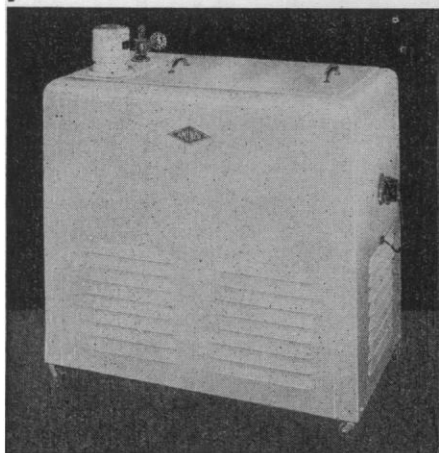
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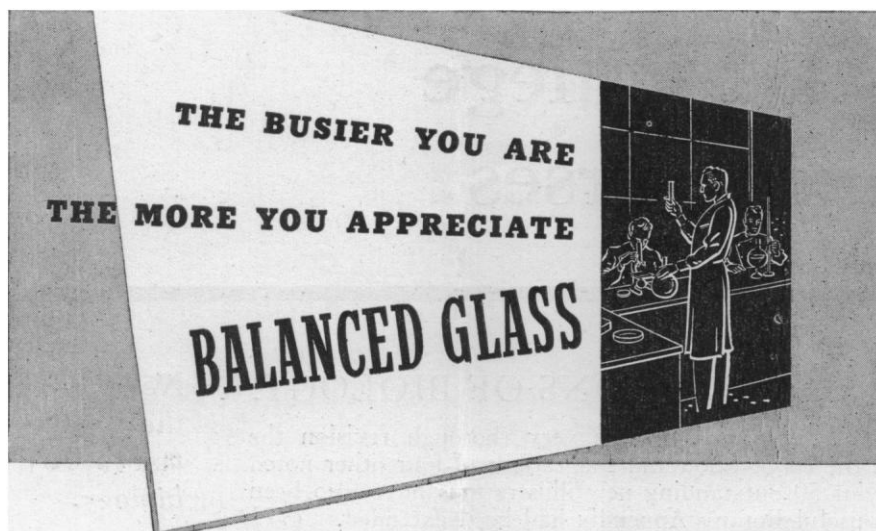
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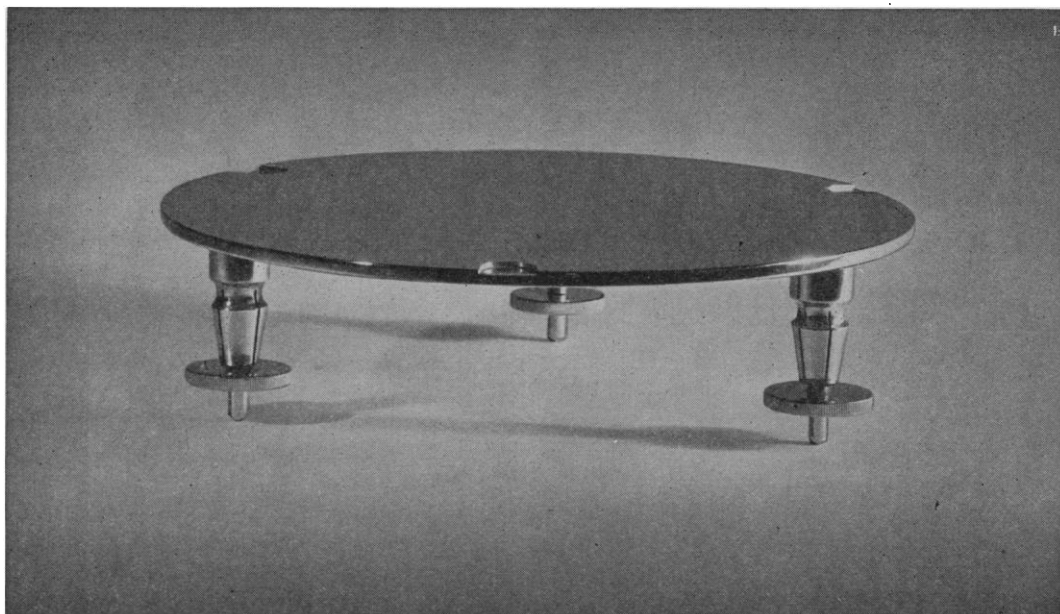
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FRIDAY, MAY 16, 1941

NO. 2420

Annual Meeting of the American Philosophical Society: PROFESSOR EDWIN G. CONKLIN 461

Obituary:

Cary LeRoy Hill: EDWARD C. CRAFTS. *Recent Deaths* 466

Scientific Events:

The New York Zoological Park; The Chicago Meeting of the American Society for Testing Materials; One Hundred and Twentieth Anniversary of the Philadelphia College of Pharmacy and Science; The Ninth Summer Conference on Spectroscopy and Its Applications; Conference on Molecular Structure; The Washington Meeting of the American Geophysical Union 467

Scientific Notes and News 470

Discussion:

Heavy Carbon Production by Thermal Diffusion: PROFESSOR WILLIAM W. WATSON. *A Giant Rodent from the Oligocene:* DR. GEORGE GAYLORD SIMPSON. *The Velocity of Light:* DR. ARTHUR BESSEY SMITH 473

Scientific Books:

Biological Oxidations: DR. KURT G. STERN 475

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or Inactivation of Biotin: DR. PAUL GYÖRGY and OTHERS. *The Excretion of Injected Heparin in the Urine of Mice and Dogs:* DR. ALFRED LEWIN COPLEY. *The Artificial Synthesis of a 42-Chromosome Wheat:* E. J. BRITTEN and PROFESSOR W. P. THOMPSON 477

Scientific Apparatus and Laboratory Methods:

An Antimony Electrode for the Continuous Recording of the Acidity of Human Gastric Contents: DR. HOWARD W. HAGGARD and DR. LEON A. GREENBERG 479

Science News 8

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ANNUAL MEETING OF THE AMERICAN PHILOSOPHICAL SOCIETY

By Professor EDWIN G. CONKLIN

VICE-PRESIDENT AND EXECUTIVE OFFICER; PROFESSOR EMERITUS OF BIOLOGY,
PRINCETON UNIVERSITY

THE American Philosophical Society held its annual meeting in its hall on Independence Square, Philadelphia, on Thursday, Friday and Saturday, April 24, 25, 26. The annual meeting of the council was held on the afternoon and evening preceding the session on Thursday. This was an innovation intended to give more time to the important work of the council than was possible heretofore when it met only at dinner on the evening preceding the executive session. President and Mrs. Morris entertained the council at dinner, which was given this time in the society's hall.

The program on Thursday morning and afternoon included thirteen papers in the physical and biological sciences, abstracts of which are printed below. On Thursday evening, Dr. Hugh S. Taylor, professor of

chemistry at Princeton University, gave the Franklin Medal Lecture on "Large Molecules through Atomic Spectacles." The lecture was illustrated by models, slides and moving pictures showing novel conceptions of the structure of large molecules. The Benjamin Franklin Medal of the society was presented to Professor Taylor at the close of his lecture.

In the old days congenial members of this society used to adjourn to a Rathskeller or beer garden after an evening lecture, where light refreshment and light or heavy discourse were in order. These aftermath meetings were often most profitable and enjoyable. With the thought of reviving as far as possible this custom, round table parties were held in the hall of the society following Professor Taylor's lecture and

The antimony electrode is extensively used in measuring pH in industrial processes,³ but it is rarely used in physiological investigations. It does not function satisfactorily in strongly alkaline solutions⁴ or in the presence of powerful oxidizing agents⁵ as in the blood, but these limitations are not encountered in dealing with gastric contents. The antimony electrode is less sensitive than the glass electrode,⁴ but is accurate to a change of 0.1 in pH, which is adequate for measurements on gastric contents. The greatest disadvantage of the antimony electrode has been the

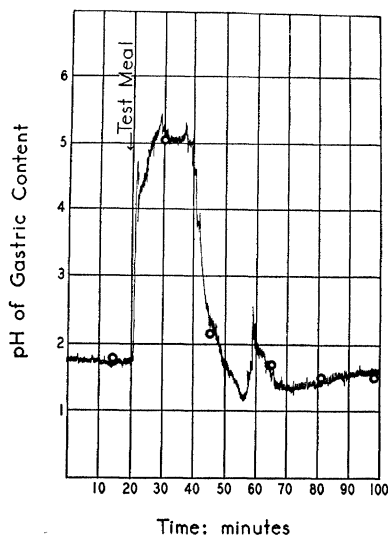


FIG. 1. Continuous record of pH of gastric contents obtained with antimony electrode before and after test meal. Dots represent values obtained with glass electrode on gastric contents withdrawn with Levin tube.

tendency for its current to drift and thus yield erratic and erroneous values.^{6,7} Many attempts have been made to overcome this defect by obtaining a surface coating of stable oxides.⁵ We have succeeded in preparing an electrode which is stable in gastric contents by treating the metal with bromine water.

In preparation for use, the electrode is polished with fine emery paper and immersed for 30 minutes in a 1 per cent. solution of bromine in water; it is then washed in water and dried with a soft cloth but is not polished. At intervals of approximately one week during daily use, the electrode is polished with emery paper and treated with bromine water as de-

scribed. The electrode thus prepared responds to change in pH within the range 1.0 to 7.0 without appreciable lag and without drift.

There is no danger of any toxic action from the antimony when the electrode is swallowed. None of the subjects, in whose stomach the electrode was retained for several hours each day for nearly a month, complained of any disturbance. As further proof, an electrode was carefully weighed and placed in $n/10$ HCl; after 24 hours of immersion the loss of weight was 0.2 mg; and after 48 hours, 0.3 mg.

The electrode does not function satisfactorily for one to two hours after a full-sized meal, presumably because it may become embedded in large masses of food. This difficulty is not encountered with any of the ordinary test meals used for clinical purposes.

Fig. 1 shows a tracing from a typical record of the pH of the gastric contents before and after a test meal. In this experiment a Levin tube was also passed into the stomach and specimens of the contents withdrawn; the pH of these specimens was determined in the usual manner with the glass electrode; the values obtained are shown as dots in Fig. 1. The agreement between the values found is within pH 0.15. In the record shown here the pH of the gastric contents reached a minimum value 36 minutes after the test meal was given. An abrupt but brief rise in pH then occurred. This momentary rise is characteristic of many records we have obtained.

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- ³ G. A. Perley, *Indus. Eng. Chem.* (Analytical Ed.) 11: 316, 1939.
- ⁴ G. A. Perley, *Indus. Eng. Chem.* (Analytical Ed.) 11: 319, 1939.
- ⁵ T. R. Ball, The Antimony Electrode in pH Measurements. *Trans. Electrochem. Soc.* 72: 139, 1937.
- ⁶ E. J. Roberts and F. Fenwick, *Jour. Am. Chem. Soc.*, 50: 2125, 1928.
- ⁷ H. W. Haggard and L. A. Greenberg, *Am. Jour. Digest Dis.* (in press).

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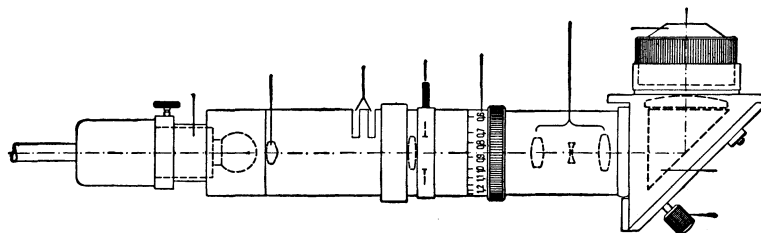
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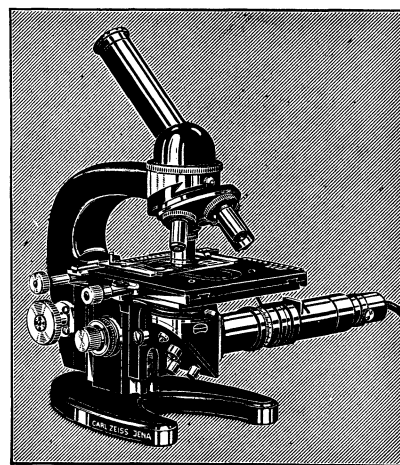
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