ments in which we tested several hundred substances on several hundred thousand pig Ascaris *in vitro,* collected the parasites passed, as well as those found at autopsy, in many hundreds of dogs infested with Ascaris, and egg-counted over 3,000 human cases of ascariasis before and after treatment.

During the course of these experiments we checked as a matter of interest, but with no intent of publication, the effect of 121 substances on both earthworms and pig Ascaris, using the *in vitro* method described in our "Methods" paper. The earthworms were tested in distilled water at room temperature, while the Ascaris were kept at 35° in saline. Of these 121 chemicals chosen from *inorganic substances, hydrocarbons, halogenated hydrocarbons, alcohols, phenols, ethers, organic acids* and their *salts, organic sulfur compounds, terpenes, sesquiterpenes, alkaloids, enzymes, glucosides, plant products, dyes, substances liberating oxygen* or *iodine,* as well as a miscellaneous group, only 6 per cent. showed a fairly close correspondence of anthelmintic action on earthworms and pig Ascaris. In 58 per cent. the Ascaris were alive after twenty hours, while the earthworms had died in from two minutes to six hours. Sixty-seven per cent. killed earthworms in thirty minutes or less, while only 8 per cent. of these substances killed Ascaris in that time.

Conclusion

(1) The use of earthworms as a test object for evaluating the activity of anthelmintics to be used in human intestinal helminth infestation is irrational.

(2) A comparative study of the lethality of 121 widely diversified chemical substances on both earthworms and pig Ascaris shows no correlation of action.

(3) *In vitro* tests of human ascaricides on pig Ascaris, which is morphologically indistinguishable from human Ascaris, are of value.

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CATALYSIS OF FORMALDEHYDE TO REDUCING SUGARS BY ASCORBIC ACID

Kusin[1] has recently shown that compounds capable of forming en-diol groups in alkaline solution (glucose, fructose, benzoin, etc.) markedly catalyze the production of reducing sugars from formaldehyde in the presence of calcium hydroxide. According to Haworth's formula ascorbic acid contains the en-diol group and, if Kusin's theory is correct, might be expected to catalyze the reaction. The writers have investigated this possibility and have found that ascorbic acid is a very active catalyst. Table 1 summarizes the results of experiments in which the rate of formation of reducing substances was followed in mixtures of formaldehyde and calcium hydroxide both with and without the addition of ascorbic acid. Ascorbic acid and excess formaldehyde were removed from the solutions before sugar estimations.

TABLE 1

Reactants	Time Minutes	Red. Subs. as glucose mg. per 100 cc
Formaldehyde + ascorbic acid + $Ca(OH)_2$	0	24.
"	30	220.
"	60	391.
"	90	312.
"	120	269.
"	180	199.
Formaldehyde + $Ca(OH)_2$	0	0.
"	60	0.
"	120	11.
"	180	154.
"	240	94.

The disappearance of formaldehyde from the solutions was about complete at the time of maximum reduction.

It will be noted that in the reaction catalyzed by ascorbic acid maximum reduction was attained before the appearance of reducing substances in the control. In both cases the curves show a maximum after which the reducing values decrease. This is possibly due to higher reducing equivalents of the smaller molecules formed early in the reaction followed by polymerization of these to larger molecules of lower reducing values.

Kusin has isolated an intermediate compound of benzoin and formaldehyde from reaction in an alkaline solution. We have obtained evidence that ascorbic acid combines with formaldehyde even in an acid solution. The suggestion has been made that simple sugars play a catalytic rôle in the photosynthetic production of carbohydrates from formaldehyde in green leaves. The evidence presented here suggests that ascorbic acid may be an active photosynthetic catalyst.

Further work is in progress and this, with experimental details, will be published elsewhere.

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THE CONDENSATION OF β-CYCLOCITRAL WITH DIMETHYLACROLEIN

According to the principle of vinylogy[1] the condensation of β-cyclocitral (I) with two moles of

[1] A. Kusin, *Ber.* 68: 619, 1494 and 2169 (1935).

[1] Fuson, *Chem. Rev.,* 16: 1, 1935.