not be prepared. A large number of inquiries are received for chemicals which we could never hope to furnish; in some instances, the preparations could be undertaken, but it is questionable whether the time devoted to working out the method and preparing a stock might not be better applied to some problem for which there is greater urgency. Our desire is to serve the research chemists of the United States, but to do this to best advantage it is necessary to consider the interest of the greatest number. We acknowledge with gratitude the continued support of the chemical manufacturers, who have supplied us not only with their regular products, but often with those which are available in quantities too small to place on the open market. The amount of chemicals sold continues to increase slowly but steadily, and the department is now almost self-supporting. It is at present being transferred to new laboratories especially designed and erected for the work, and it is expected that greater efficiency will be possible than in the improvised laboratory where the work was begun.

The production of benzoic acid and benzophenone from benzene and phosgene: Robert E. Wilson and Everett W. Fuller.

The nature of the reactions of anilines upon nitrosophenol: Carleton E. Curran and C. E. Boord. Experimental evidence shows that the first reaction product between aniline and nitrosophenol is quinone phenylhydrazone. Dilution or neutralization of the reaction mixture converts this substance into its tantomer phenyl-azophenol. Subsequent action of aniline upon the quinone-phenylhydrazone converts it into mono-anilino quinonephenylhydrazone, dianilino quinone and azophenine, in turn. The theory is proposed that the formation of indamines by the action of anilines upon nitrosophenol is due to the semidine rearrangement of quinone-phenylhydrazones.

Reduction of polynitrophenols by hydrogen sulphide in the presence of ammonia: L. Chas. Raiford. In the preparation of starting material with which to test further the migration of acyl noted in a previous paper (*Jour. Am. Chem. Soc.*, 41, 2068 (1919)), with a view to determining the effect of acid-forming substituents in the aminophenol 2, 4-dinitrophenol was reduced by hydrogen sulphide in the presence of ammonia in the usual way. Contrary to what has heretofore been reported, isomeric substances were obtained. Work is in progress to determine the effect of other substituents (compare Anschutz und Heusler, *Ber.*, 19, 2161 (1886)).

The action of ammonia and substituted amines on allophanic ester: F. B. Dains and E. Wertheim.

Hydrazoisopropane: H. L. Lochte and J. R. Bailey.

A convenient method for preparing certain bromohydrins: J. B. Conant and E. L. Jackson.

Addition reactions involving an increase in vaence of a single atom: J. B. Conant.

New derivations of thymol and carvacrol: D. S. L. Sherk and Edward Kremers. The quinhydrone hypothesis of plant pigments, as it grew out of the biochemistry of the Monardas, necessitated a revision of the underlying compounds. This study has been continued, especially along the line of intramolecular changes such as manifest themselves in connection with the nitroso compounds of the above mentioned phenols and their isonitroso rearrangement products.

The action of amines upon thymoquinone: Nellie A. Wakeman and Harlan G. Groffman. Dimethylamidothymoquinone, prepared according to Zincke, yields a platinic chloride double salt containing 41 per cent. of platinum, corresponding to the union of one molecule of the base with two of acid platinic chloride. Thymoquinone treated with benzylamine, in alcoholic solution, yields dibenzylaminothymoquinone, with some mono-benzylaminothymoquinone. Thymoquinone with aniline, also with p-toluidine, under the same conditions, yields the di- derivative. No mono- derivatives have been isolated here. Thymoquinone with piperidine, under the same conditions, yields a pale purple crystalline derivative, the constitution of which has not yet been determined.

Organic mercury compounds of phenol: Frank C. Whitmore and E. B. Middleton.

Charles L. Parsons,
Secretary

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