a basic salt, the high atomic weight $R^{iv} = 236.3$. But I stated expressly, and I feel obliged to repeat it, that these fractions show a great tendency to form basic salts. Assuming these to be normal, a higher atomic weight than the true one is obtained. This is true especially in regard to the oxalate.

The splitting up of thorium into Th^{α} and Th^{β} was, of course, not so sensational an event as the announcement from America of the splitting up of thorium into 'carolinium' and 'berzelium.'

BOHUSLAV BRAUNER.

Bohemian University, Prague,
April 18.

Those who have read my work and heard my recent paper delivered before the Washington, New York and North Carolina sections of the American Chemical Society do not require further information regarding the above. In view of the fact that many British men of science are not familiar with the work and may be misled, it has been deemed wise to despatch the following to the editor of *Nature*.

Re Thorium.—The elementary nature of thorium has been questioned by several workers, namely, Chroustschoff in 1889 (*J. russ. phys. Chem. Ges.*, 29, 206), Rutherford in 1899 (*Phil. Mag.*, 49, 2, 1900), Crookes in 1900 (*Proc. Roy. Soc.*, 66, 406) and in 1901 Brauner (*Proc. Chem. Soc.*, 17, 67) and Baskerville working independently (*Journ. Am. Chem. Soc.*, 23, 761). The methods employed were different in each case.

The undersigned has made no claim of priority as to the idea of the complexity of thorium, but he distinctly claims to have applied novel methods and an old one, which demonstrate to the satisfaction of himself and others familiar with the work, not only the complexity of old thorium, but the existence of two new elements to which the names of carolinium and berzelium have properly been given. The old method was used by Berzelius, who died thirty years before the plaintiff, according to his own statement (April 28, p. 606), began his work on the separation of the rare earths.

Scientific men will await the appearance of the paper, which will be published shortly in the *Journal of the American Chemical Society*, and see that all workers have received full credit for their share in the solution of the question. In the meantime, the letter adverted to, carrying much that is true and a distortion, which any one may verify by reference to the literature, to say the least is in poor taste.

For fear lest the old proverb, '*qui tacet consentire videtur*,' carry too much influence, the above statement is reluctantly made.

CHAS. BASKERVILLE.

UNIVERSITY OF NORTH CAROLINA, U. S. A.,
May 17, 1904.

A REDDISH-BROWN SNOWFALL.

TO THE EDITOR OF SCIENCE: An incident which should, perhaps, be recorded is that of a reddish-brown snowfall which occurred at this place on February 2 last (1904). A light snow was falling on that day and about noon the character of the snow-fall changed to a reddish-brown or light chocolate color. This continued for half or three quarters of an hour, after which the snow-fall of ordinary appearance continued during the afternoon, the colored snow appearing as a well-defined layer between the white snow which fell before and after it. An examination under the microscope showed numerous irregular-shaped, semi-transparent particles with an appearance similar to feldspar. Nitric and muriatic acid applied to them gave no apparent result. Examined microscopically during the snow-fall it appeared that the particles were not carried on the snow, but were embedded in the snow crystals. Other ordinary contaminations were present, but were plainly distinguishable from the peculiar particles in the snow crystals. The phenomenon was observed in two or three near-by towns, but, so far as learned, not outside this immediate vicinity.

EDWARD LINDSEY.

WARREN, PA.,

SPECIAL ARTICLES.

MENTAL EFFICIENCY AND HEALTH.

IN the address as president of the American Society of Naturalists, read by Professor Cattell at the annual dinner, January 1, 1903, and printed in this journal, April 10, 1903, is inserted a table giving the grades for different